

## 729. The Crystal and Molecular Structure of Triphenylphosphorus.

By J. J. DALY.

The structure of triphenylphosphorus, PPh<sub>3</sub>, has been determined by single-crystal X-ray diffraction methods. A three-dimensional least-squares refinement has been carried out. The molecule has the expected pyramidal configuration about the phosphorus atom but, owing to the unequal rotation of the benzene rings about the P-C bonds, possesses no symmetry. The mean P-C bond length is 1.828 Å; none of these three bonds differs significantly from the mean.

THE determination of the correct structure of triphenylphosphorus in the solid state was reported recently.<sup>1</sup> The parameters obtained from this two-dimensional X-ray analysis have now been refined by the least-squares method, using three-dimensional data recorded photographically from Cu K<sub>α</sub> radiation. The unit cell, space group, and density have been known for some time and the data have been summarised.<sup>2</sup>

### EXPERIMENTAL AND RESULTS

C<sub>18</sub>H<sub>15</sub>P, *M* = 262.3, Monoclinic, *a* = 11.413, *b* = 15.032, *c* = 8.500 Å, β = 92° 53', *U* = 1456.4 Å<sup>3</sup>, *Z* = 4, *D*<sub>c</sub> = 1.196 g./cc. Space group *P*2<sub>1</sub>/*a*(*C*<sub>2h</sub><sup>5</sup>). Cu K<sub>α</sub> radiation, single-crystal oscillation and Weissenberg photographs. The cell constants were determined from high-angle reflexions on oscillation photographs with asymmetric film mounting and are estimated to be

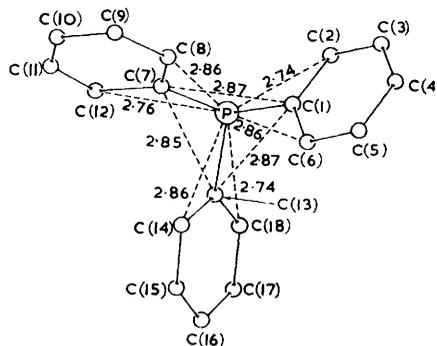


FIG. 1. The *a*-axis projection of the molecule, showing the numbering of the atoms and intramolecular distances less than 3.0 Å between non-bonded atoms, excluding C...C in benzene rings.

accurate to 1 part in 700. The cell dimensions are in reasonable agreement with previous values.<sup>2,3</sup> The axial ratio *a* : *b* = 0.7592 agrees with the value found by optical goniometry<sup>4</sup> (0.7590).

Multiple-film Weissenberg photographs were taken round all three principal axes. The relative intensities of the diffracted beams were estimated visually by comparison with a calibrated strip. All layers were brought to the same relative scale by correlation through common reflexions. 2290 reflexions were observed to be non-zero. Systematic absences were: *h*0*l* absent if *h* odd, 0*k*0 absent if *k* odd. Thus, the space group is *P*2<sub>1</sub>/*a*. No correction has been made for absorption ( $\mu = 15.7 \text{ cm.}^{-1}$ ).

All three-dimensional calculations were done on an Elliott 803B computer using a programming system devised by Daly, Stephens, and Wheatley.<sup>5</sup> The quantity minimised in the least-squares refinement was  $\sum w(|F_o| - |F_c|)^2$ . The anisotropic thermal parameters were used in the

<sup>1</sup> Daly, *Z. Krist.*, 1963, **118**, 332.

<sup>2</sup> *Structure Reports*, 1950, **13**, 551.

<sup>3</sup> Howells, Lovell, Rogers, and Wilson, *Acta Cryst.*, 1954, **7**, 298.

<sup>4</sup> Porter and Spiller, "Baker Index of Crystals," 1956, Vol. II, Part 3, M 2692.

<sup>5</sup> Daly, Stephens, and Wheatley, Monsanto Research S.A. Final Report No. 52 (1963).

TABLE 1.

Co-ordinates, standard deviations, and orthogonal co-ordinations in Å.

	<i>X</i>	<i>Y</i>	<i>Z</i>	<i>X'</i>	<i>Y'</i>	<i>Z'</i>
P	2.4259(11)	0.5627(12)	1.2005(11)	2.3655	0.5627	1.1990
C(1)	3.5649(44)	1.9779(47)	1.4101(48)	3.4940	1.9779	1.4083
C(2)	3.3675(52)	2.8035(52)	2.5224(54)	3.2406	2.8035	2.5192
C(3)	4.1901(62)	3.9015(54)	2.7772(65)	4.0504	3.9015	2.7737
C(4)	5.1890(64)	4.2199(59)	1.9302(69)	5.0919	4.2199	1.9277
C(5)	5.4083(59)	3.3982(63)	0.8255(68)	5.3668	3.3982	0.8245
C(6)	4.5845(51)	2.2879(54)	0.5549(57)	4.5566	2.2879	0.5542
C(7)	3.4373(44)	-0.8618(48)	1.8018(45)	3.3467	-0.8618	1.7995
C(8)	4.7024(49)	-0.7426(51)	2.4490(51)	4.5792	-0.7426	2.4459
C(9)	5.3381(61)	-1.8812(58)	2.9977(63)	5.1873	-1.8812	2.9939
C(10)	4.7318(65)	-3.1393(63)	2.8811(70)	4.5869	-3.1393	2.8774
C(11)	3.4760(66)	-3.2609(62)	2.2399(73)	3.3633	-3.2609	2.2371
C(12)	2.8250(59)	-2.1321(57)	1.6851(64)	2.7402	-2.1321	1.6830
C(13)	2.3314(45)	0.2894(48)	-0.6123(46)	2.3622	0.2894	-0.6115
C(14)	3.3149(54)	-0.4223(66)	-1.2903(52)	3.3798	-0.4223	-1.2887
C(15)	3.1694(70)	-0.5393(72)	-2.6862(63)	3.3045	-0.5393	-2.6828
C(16)	2.0659(75)	0.0040(71)	-3.4012(60)	2.2370	0.0040	-3.3969
C(17)	1.0800(78)	0.6937(61)	-2.7178(64)	1.2167	0.6937	-2.7143
C(18)	1.2311(57)	0.8358(53)	-1.3122(55)	1.2971	0.8358	-1.3105

TABLE 2.

Anisotropic thermal parameters and their standard deviations in Å<sup>2</sup>.

	<i>U</i> <sub>11</sub>	<i>U</i> <sub>22</sub>	<i>U</i> <sub>33</sub>	2 <i>U</i> <sub>12</sub>	2 <i>U</i> <sub>23</sub>	2 <i>U</i> <sub>13</sub>
P	0.0504(5)	0.0626(6)	0.0584(6)	-0.0003(9)	-0.0077(10)	0.0005(9)
C(1)	0.0590(22)	0.0622(25)	0.0662(25)	0.0069(39)	-0.0100(41)	-0.0127(38)
C(2)	0.0780(30)	0.0656(30)	0.0749(29)	0.0084(47)	-0.0220(47)	-0.0173(48)
C(3)	0.0938(38)	0.0674(31)	0.0970(41)	-0.0045(55)	-0.0234(58)	-0.0326(63)
C(4)	0.0953(39)	0.0671(34)	0.1120(46)	-0.0097(59)	0.0028(62)	-0.0507(67)
C(5)	0.0801(34)	0.0819(36)	0.1100(47)	-0.0295(60)	0.0173(67)	0.0062(63)
C(6)	0.0671(26)	0.0712(30)	0.0900(35)	-0.0205(48)	-0.0001(54)	0.0125(48)
C(7)	0.0594(23)	0.0643(26)	0.0603(23)	-0.0090(39)	-0.0056(40)	-0.0020(37)
C(8)	0.0648(25)	0.0741(31)	0.0709(28)	0.0142(45)	-0.0031(47)	-0.0183(43)
C(9)	0.0895(35)	0.0762(35)	0.0943(39)	0.0248(58)	0.0093(58)	-0.0277(60)
C(10)	0.0978(40)	0.0771(36)	0.1054(45)	0.0213(63)	0.0172(66)	-0.0155(68)
C(11)	0.1019(42)	0.0660(34)	0.1200(52)	0.0080(61)	0.0119(66)	-0.0188(75)
C(12)	0.0849(34)	0.0671(31)	0.1013(41)	-0.0125(53)	-0.0030(57)	-0.0009(60)
C(13)	0.0596(23)	0.0679(28)	0.0601(22)	-0.0224(41)	-0.0063(41)	-0.0032(37)
C(14)	0.0754(32)	0.1085(42)	0.0671(29)	-0.0341(59)	-0.0471(55)	-0.0198(49)
C(15)	0.0999(43)	0.1223(52)	0.0794(35)	-0.0682(79)	-0.0472(72)	0.0351(63)
C(16)	0.1315(53)	0.1023(45)	0.0702(32)	-0.1054(86)	0.0150(66)	-0.0091(65)
C(17)	0.1360(55)	0.0822(39)	0.0830(37)	-0.0604(73)	0.0155(60)	-0.0655(74)
C(18)	0.0867(34)	0.0664(29)	0.0764(31)	-0.0248(49)	0.0096(47)	-0.0428(53)

form  $\exp - 2\pi^2 \left( \sum_{i,j=1}^3 U_{ij} h_i h_j \right)$ . The weighting scheme was the one described by Cruickshank: <sup>6</sup>  $w = 1/(a + |F_o| + c|F_o|^2)$ , with  $a = 2|F_{\min}|$  and  $c = 2/|F_{\max}|$ .

The final values of *R* and *R'* were 0.101 and 0.018, respectively. The number of planes used in the final refinement cycle was 2260. The final co-ordinates, followed by their estimated standard deviations (e.s.d.) in parentheses, and the orthogonal co-ordinates are given in Table 1. Thus, 2.4259 (11) means that the co-ordinates of the atom (2.4259 Å) has an e.s.d. of 0.0011 Å. The transformation of the monoclinic co-ordinates (*X*, *Y*, *Z*) to orthogonal co-ordinates (*X'*, *Y'*, *Z'*) is given by:  $X' = X + Z \cos \beta$ ;  $Y' = Y$ ;  $Z' = Z \sin \beta$ . The thermal parameters *U*<sub>*ij*</sub> and their e.s.d.s' are listed in Table 2. *F*<sub>o</sub>, *F*<sub>c</sub>, and Δ, from the penultimate set of parameters, are given in Table 3. The maximum parameter shift in the final cycle was 0.3 of one e.s.d. [the *Z*-co-ordinate of C(4)]. The bond lengths and angles with their e.s.d.s' are given in Table 4. Fig. 1 shows the *a*-axis projection of the molecule, the numbering of the atoms and non-bonded intramolecular distances below 3 Å, excluding carbon-carbon distances across a benzene ring. Fig. 2 shows the *a*-axis projection of the unit-cell contents, and Fig. 3 the *c*-axis projection of

<sup>6</sup> Cruickshank *et al.*, "Symposium on Computer Methods," Pergamon, London, 1961.













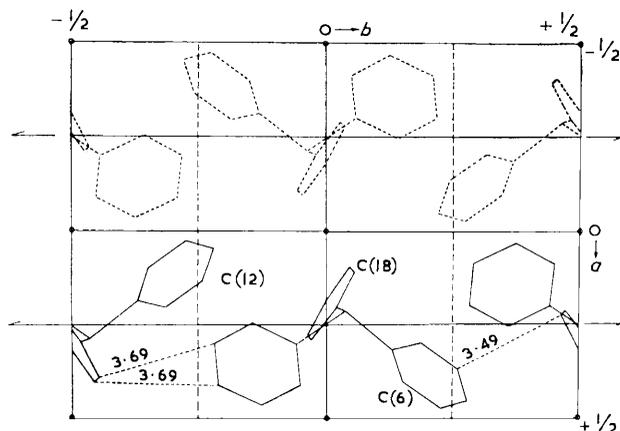
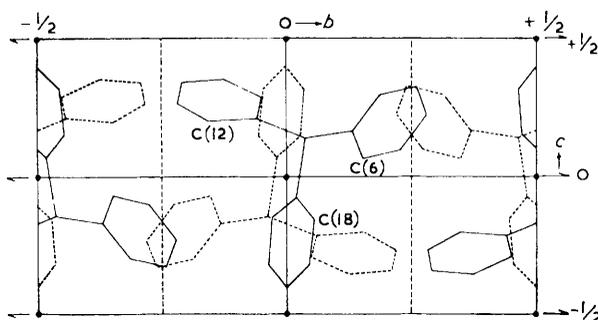




TABLE 4.

Bond lengths (Å) and angles (°), and their standard deviations.

P-C(1) .....	1.8221(48)	P-C(7) .....	1.8310(47)	P-C(13) .....	1.8310(48)
C(1)-C(2) .....	1.4071(70)	C(7)-C(8) .....	1.3968(68)	C(13)-C(14) .....	1.4144(74)
C(2)-C(3) .....	1.3878(80)	C(8)-C(9) .....	1.4023(79)	C(14)-C(15) .....	1.4011(89)
C(3)-C(4) .....	1.3790(88)	C(9)-C(10) .....	1.3989(90)	C(15)-C(16) .....	1.3945(97)
C(4)-C(5) .....	1.4029(90)	C(10)-C(11) .....	1.3863(94)	C(16)-C(17) .....	1.4080(97)
C(5)-C(6) .....	1.4008(83)	C(11)-C(12) .....	1.4034(90)	C(17)-C(18) .....	1.4133(88)
C(6)-C(1) .....	1.3981(71)	C(12)-C(7) .....	1.4124(75)	C(18)-C(13) .....	1.3862(72)
C(1)-P-C(7) .....	103.57(21)	C(2)-C(1)-P .....	115.78(36)	C(6)-C(1)-P .....	124.94(37)
C(7)-P-C(13) .....	102.07(21)	C(12)-C(7)-P .....	116.27(37)	C(8)-C(7)-P .....	123.93(36)
C(13)-P-C(1) .....	103.34(21)	C(18)-C(13)-P .....	116.18(37)	C(14)-C(13)-P .....	123.17(38)
C(1)-C(2)-C(3) ...	120.26(49)	C(7)-C(8)-C(9) ...	119.62(47)	C(13)-C(14)-C(15) .....	118.67(53)
C(2)-C(3)-C(4) ...	120.72(56)	C(8)-C(9)-C(10) ...	120.77(55)	C(14)-C(15)-C(16) .....	121.21(61)
C(3)-C(4)-C(5) ...	119.68(58)	C(9)-C(10)-C(11) .....	119.75(60)	C(15)-C(16)-C(17) .....	119.82(63)
C(4)-C(5)-C(6) ...	120.18(57)	C(10)-C(11)-C(12) .....	120.24(60)	C(16)-C(17)-C(18) .....	119.31(60)
C(5)-C(6)-C(1) ...	119.83(50)	C(11)-C(12)-C(7) ...	120.01(54)	C(17)-C(18)-C(13) .....	120.33(52)
C(6)-C(1)-C(2) ...	119.28(45)	C(12)-C(7)-C(8) ...	119.59(45)	C(18)-C(13)-C(14) .....	120.64(46)

FIG. 2. *a*-Axis projection of the unit-cell contents.FIG. 3. *c*-Axis projection of the unit-cell contents and van der Waals contacts less than 3.7 Å

The bond lengths have not been corrected for torsional oscillation; any such correction, despite the large values of the  $U_{ij}$ 's, can only worsen the agreement between the mean C-C bond length and the standard value of 1.397 Å.

## DISCUSSION

The rules used here to compare experimentally determined quantities follow those given by Cruickshank and Robertson.<sup>7</sup> All the carbon atoms lie within  $1.7\sigma_{\text{rms}}$  of the appropriate

<sup>7</sup> Cruickshank and Robertson, *Acta Cryst.*, 1953, **6**, 698.

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least-squares planes;  $\sigma_{\text{rms}}^2 = [\sigma^2(X) + \sigma^2(Y) + \sigma^2(Z)]/3$ . The mean bond lengths and angles in the benzene rings are:

Ring	A	B	C	All
Mean C-C length .....	1.3960(31)	1.4000(31)	1.4029(35)	1.3996(20) Å
Mean C-C-C angle .....	119.99(22)	120.00(22)	120.00(23)	120.00(13) Å

The most significant deviation from the mean C-C bond length is 0.0148 Å ( $\Delta/\sigma = 1.92$ ) for C(13)-C(14); the most significant deviation from the mean C-C-C angle is 1.33° ( $\Delta/\sigma = 2.42$ ) for C(13)-C(14)-C(15). The mean values of the aromatic C-C lengths agree well with standard values.<sup>8</sup> The angles are close to the theoretical value of 120°. The phosphorus atom displacements from the planes of the benzene rings are: 0.0017 Å from ring A, -0.1713 Å from ring B, and 0.0408 Å from ring C.

Although the three P-C bond lengths do not differ significantly from one another, one of the C-P-C angles [C(7)-P-C(13)] is considerably smaller than the other two:  $\Delta/\sigma = 6.00$  and 5.08. Since this angle does not involve the P-C(1) bond, and since  $C(1) \cdots C(7) = C(1) \cdots C(13) \neq C(7) \cdots C(13)$ , the symmetry of the P C(1) C(7) C(13) pyramid is closer to *m* than to *3m*. The benzene rings adopt a configuration in which the molecule possesses no symmetry. This is best seen by comparing the angles that the normals to the rings A, B, and C make with the lines P-C(13) (117.8°), P-C(1) (85.6°), and P-C(7) (115.2°), respectively. These angles should all be equal if the symmetry of the molecule is to be *3*. The B-ring is therefore twisted about 30° from the position which would give the molecule a threefold symmetry axis. This distortion cannot be due to intramolecular repulsions of the CH groups as twisting the B-ring to the symmetrical position would only reproduce contacts already present. Intermolecular repulsions may be the cause of the twisting, although the shortest of these contacts (3.49 Å) does not involve the B-ring. The shortest non-bonded intramolecular distances between atoms which are not bonded to a common atom are: 3.11 Å [C(1)  $\cdots$  C(8)], 3.19 Å [C(6)  $\cdots$  C(13)], 3.12 Å [C(7)  $\cdots$  C(14)].

It has been postulated, from the vibrational spectrum of the melt, that the benzene rings in PPh<sub>3</sub> are perpendicular to the base of the phosphorus pyramid.<sup>9</sup> In the solid state the angles between the normals to the rings A, B, and C and the normal to the pyramid base (0.90229X' - 0.10443Y' - 0.41830Z' = 2.35695) are 68.0, 36.1, and 65.3°, respectively. In a personal communication Professor Steger points out that the vibrational spectrum might be explained by assuming free rotation of the phenyl rings about the P-C bonds.

The difference between the mean P-C bond length found here (1.828 ± 0.003 Å) and the P-C length in P(CH<sub>3</sub>)<sub>3</sub> (1.841 ± 0.003 Å from microwave spectra<sup>10</sup> and 1.846 ± 0.003 Å from electron diffraction<sup>11</sup> with  $\Delta/\sigma = 3.25$  and 4.5, respectively) indicates only a slight degree of conjugation in the phenyl derivative. The mean P-C bond length in PPh<sub>3</sub> is not significantly different from the P-Ph bond in *trans*-dimesitylbis(diethylphenylphosphine)cobalt(II) (1.835 ± 0.017 Å).<sup>12</sup>

I thank Herr A. Widmer who estimated the intensities, measured the cell constants, and did much of the calculation involved in the two-dimensional work.

MONSANTO RESEARCH S.A., BINZSTRASSE 39,  
ZURICH 3/45, SWITZERLAND.

[Received, December 4th, 1963.]

<sup>8</sup> "Interatomic Distances," ed. L. E. Sutton, *Chem. Soc. Special Publ.* No. 11, 1958.

<sup>9</sup> Steger and Stopperka, *Chem. Ber.*, 1961, **94**, 3023.

<sup>10</sup> Lide and Mann, *J. Chem. Phys.*, 1958, **29**, 914.

<sup>11</sup> Bartell and Brockway, *J. Chem. Phys.*, 1960, **32**, 512.

<sup>12</sup> Owston and Rowe, *J.*, 1963, 3411.